AMENDMENTS TO DRAWING FIGURES

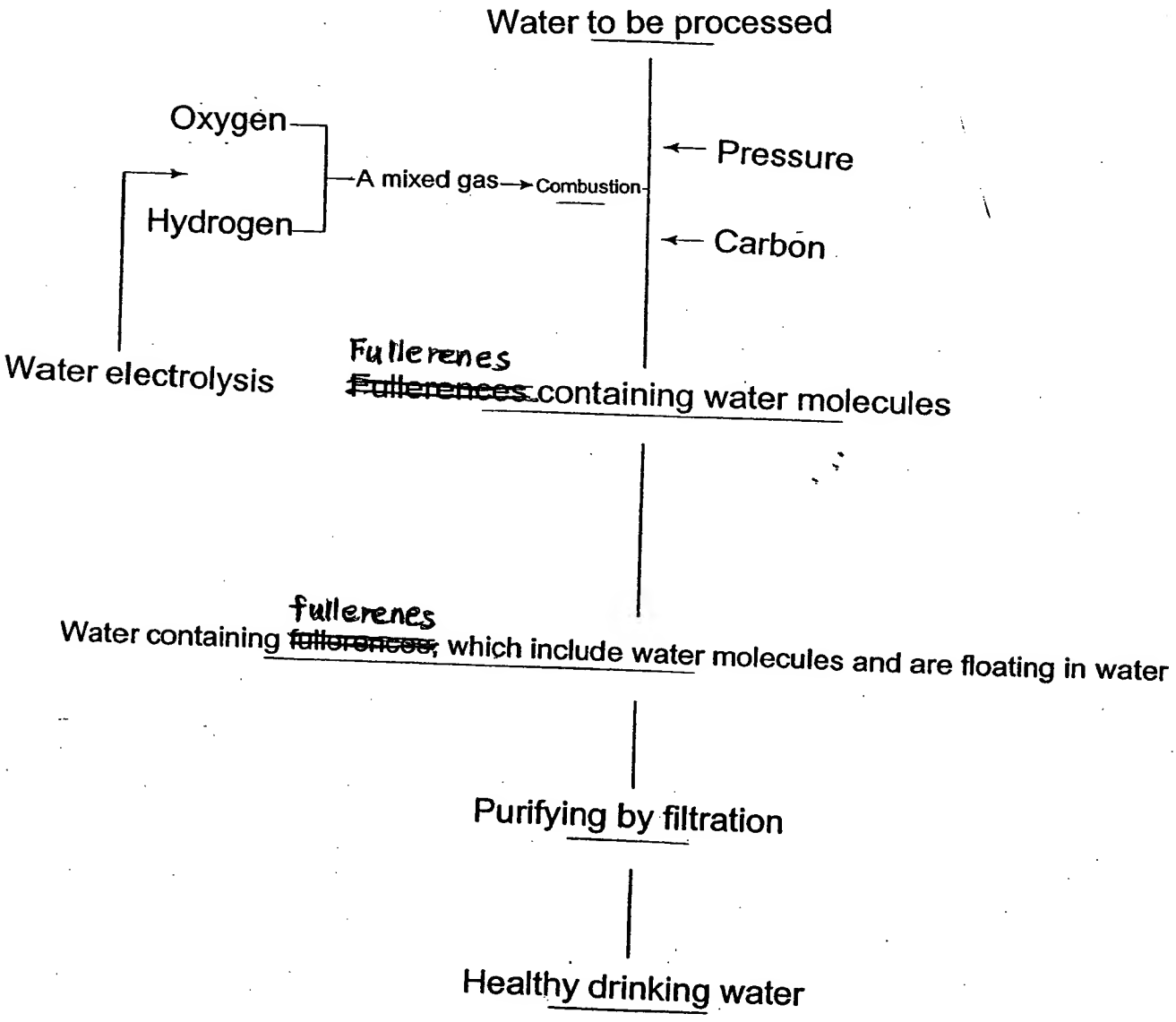
As stated below, Applicants have requested to amend Figs. 1, 3, and 4 to correct clerical errors, as shown in the attached copy of the figures wherein the corrections are indicated with red ink. That is, the term "fullerences" has been corrected to "fullerenes" in the figures.

REQUEST FOR APPROVAL OF DRAWING CHANGES

The amendment requested herein is to correct clerical errors of Figs. 1, 3, and 4 as shown with red ink in the attached copy of the figures. That is, the term "fullerenes" was misspelled as "fullerences" in the figures. No new matter has been added. Approval of this amendment is respectfully requested.

FIG. 1

A flow chart of producing water containing ~~fullerences~~ fullerenes



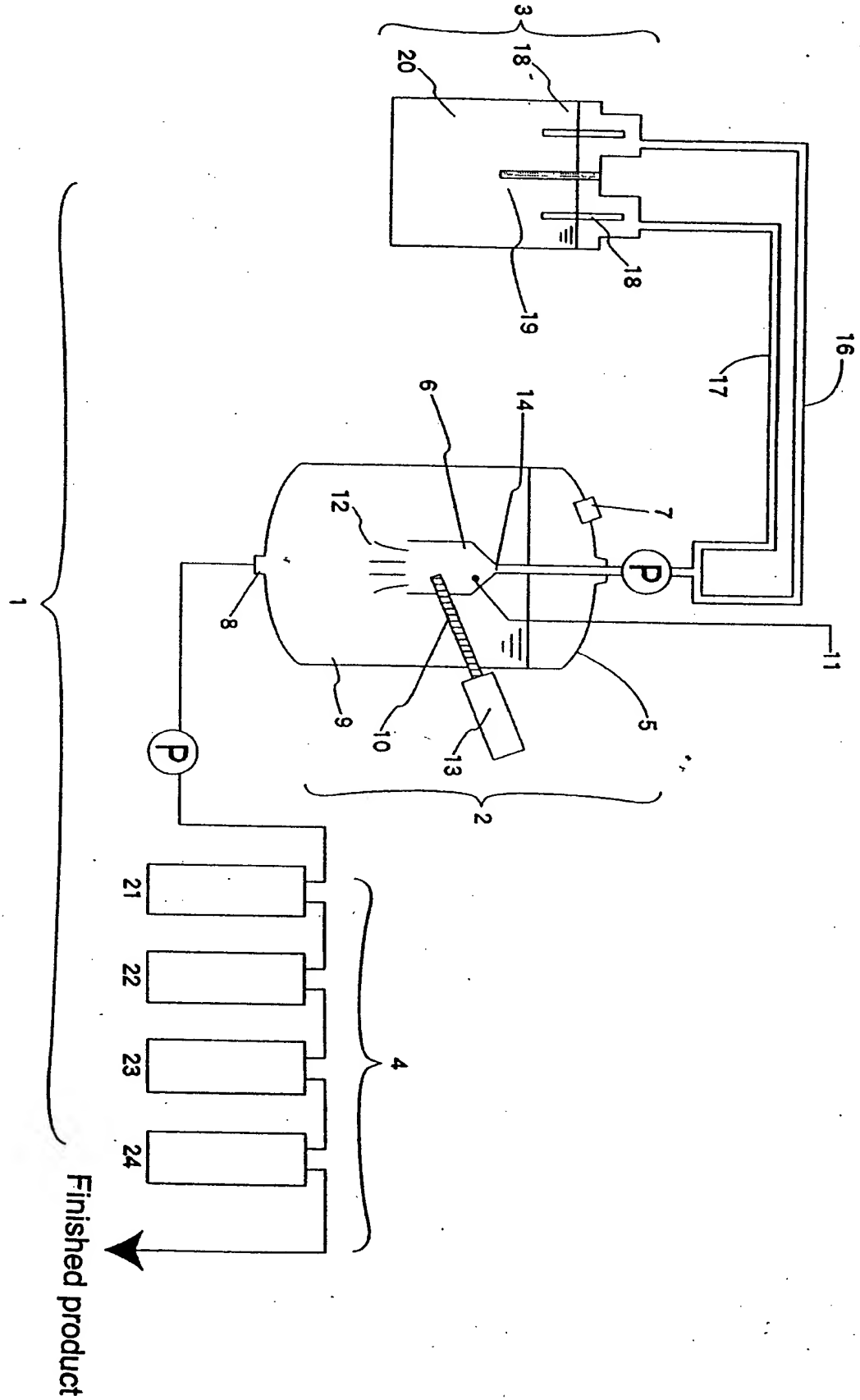


FIG. 2

FIG. 3

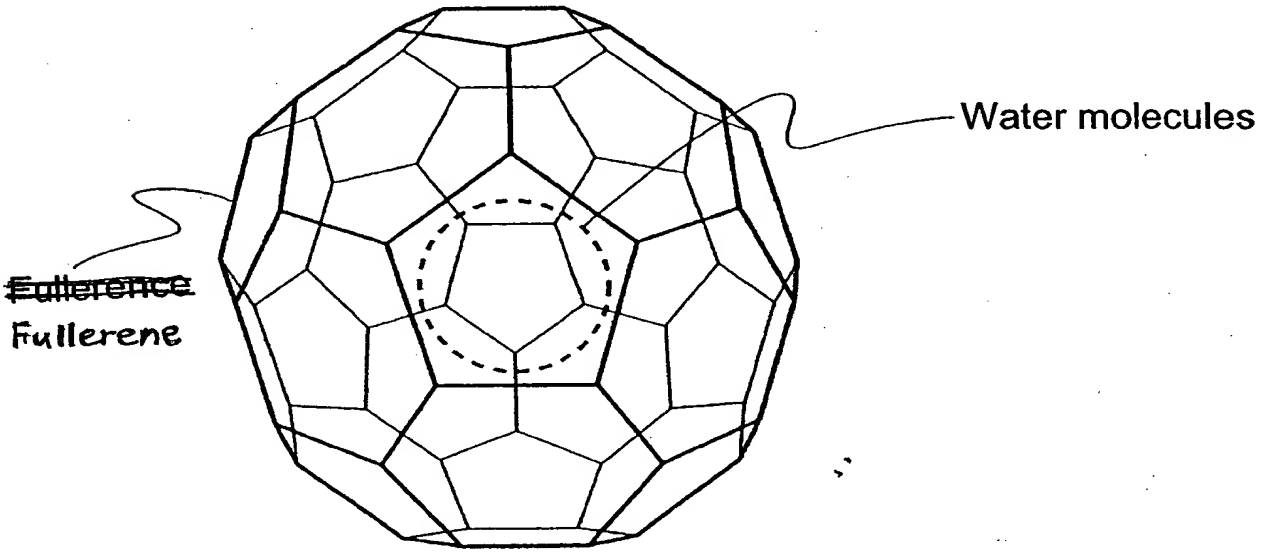


FIG. 4

Various Properties of C_{60} (prepared based on a table from Chemistry, 46, 830, 1990)

Properties (Physical Quantity)	Measured Value, etc.
·Molecular weight:	720.66
·No. of molecules:	720
·Molecular structure:	Frustum icosahedron (1_n), Diameter: ~7.1A C-C bond shared by two six-membered rings 1.391A C-C bond forming a five-membered ring 1.455A
·^{13}C-NMR spectrum (C_6D_6	δ = 143.27ppm
·Infrared adsorption spectrum (KBr pellet)/cm^{-1}	527.4, 576.4, 1182.4, 1428.5
·Infrared emission spectrum (vapor-phase, 850±100°C)/cm^{-1}	527.1, 570.3, 1169.1, 1406.9
·Raman spectrum (thin film)/cm^{-1}	273(s), 437(m), 496(s), 710(m), 774(m), 1099(w), 1250(w), 1428(m), 1470(vs), 1575(m)
·Visible ultraviolet spectrum (hexane solution, log ε in parentheses)/nm:	211(5.11), 227(sh,4.91), 256(5.24), 328(4.71), 390(3.52), 403(3.48), 492(sh,2.72), 540(2.85),568(2.78), 590(2.86), 598(2.87), 620(2.60)
·Fluorescence spectrum (toluene solution, at room temp.)/nm	No observation (thin film, 20K), 706.7(main), 787.4, 877(sh)
·Triplet energy (toluene solution)	1.56±0.03 eV (8.60±0.14 kJ/mol)
·Ionization potential	7.61±0.02 eV
·Electron affinity:	2.65±0.02 eV
·Reduction potential ($E^{1/2}$ vs Fc/Fc^+), acetonitrile/toluene, (Et_4N) BF(illegible), -10°C:	-0.98, -1.37, -1.87, -2.35, -2.85, -3.26 (V)
·Crystal structure:	Simple cubic system (249K or less) $Pa3$, Z=4, a=14.041χ (5K) Face-centered cubic system (249K or more) $Fm3$, Z=4, a=14.17±0.01χ (300K) Distance between the center of adjacent molecules: ~10.0χ
·Density:	1.729 g/cm^{-9} (5K, calculated value) 1.682 g/cm^{-9} (300K, calculated value)
·Compressibility (0~20GPa):	(5.5±0.5)x10^{-2}GPa^{-1}
·Melting point:	>700°C
·Heat of transition (249K):	~4.83kJ/mol
·heat of sublimation:	9.58±0.31 kJ/mol
·Conductivity (at room temp.):	<10.$^{-9}$Scm^{-1}
·Molar magnetic susceptibility	—(260±20)x10^{-6}emu/mol
·Transition temp. of superconducting salt Tc/K:	K_3C_{60}(18), Rb_3C_{60}(28,30), Rb_2CsC_{60}(31), $RbCs_2C_{60}$(33), K_2CsC_{60}(24), Na_2CsC_{60}(12), Na_2RbC_{60}(s.5), Na_2KC_{60}(2.5), Li_2CsC_{60}(12), Ca_xC_{60}(8.4), Sn_xC_{60}(12)
·Curie temp. of ferromagnetic salt:	$TDAE_{0.55}C_{60}$. 16.1K

* Curie temperature: Temperature at which a paramagnetic substance changes to a ferromagnetic substance when it is cooling down.
TDAE indicates tetrakis(dimethyllamino)ethylene.

(Source: K. Tanigaki & others, ~~Fullerenes~~ Fullerene, Sangyo-tosho, Oct. 27, 1992, P.16)